

Observation Sheet: pH Titration Curves

Stock concentration of NaOH: **0.0999 mol/L**

Volume of HCl: **20.00 mL**

A. Titration with the pH meter (one run shared between two students)

Note with an asterisk the measured pH and the volume of NaOH dispensed at the moment the indicator changes color.

NaOH reading (mL)	Measured pH	NaOH reading (mL)	Measured pH	NaOH reading (mL)	Measured pH
0.00	1.91	18.33	8.33		
2.05	1.92	18.41	8.87		
4.04	1.96	18.78	9.73		
6.08	2.00	20.05	10.60		
7.99	2.05	22.12	10.85		
9.94	2.10	24.06	10.98		
11.96	2.19	26.07	11.06		
13.88	2.29	28.06	11.14		
15.95	2.51	30.01	11.20		
17.26	2.89	31.93	11.23		
17.66	3.22				
17.83	3.38				
17.98	3.85				
18.07	5.25				
18.23	6.26				

Student: (name) _____ (signature) _____

Instructor: (signature) _____ **Bench #** _____ **Laboratory Section:** _____

B. Indicator titration of vinegarTitrant: **NaOH (aq)**Stock titrant concentration: **0.0999 mol/L**Dilution: The **analyte (vinegar)** was diluted: **40.00 mL** to **100.00 mL**

	run 1	run 2	run 3 (if required)
Final burette reading (mL)	34.60	39.38	
Initial burette reading (mL)	0.31	5.12	
Volume added (mL)			
Color and intensity of end point	light pink	light pink	

Student: (name) _____

(signature) _____

Instructor: (signature) _____**Bench #** _____**Laboratory Section:** _____